SYNTHESIS OF NATURAL AND UNNATURAL FRONTALIN

Takahiro HOSOKAWA, * Yoshiki MAKABE, Toru SHINOHARA, and Shun-Ichi MURAHASHI*

Department of Chemistry, Faculty of Engineering Science, Osaka University, Machikaneyama, Toyonaka, Osaka 560

(S)-(-)- and (R)-(+)-Frontalin were synthesized, respectively, in high enantiomer excess from $\beta-$ methallyl alcohol (2) by three-step procedure; (i) asymmetric epoxidation of (2), (ii) the epoxide opening with homoallylmagnesium bromide, and (iii) Pd(II)-catalyzed intramolecular acetalization.

(S)-(-)-Frontalin (1) is known to be the aggregation pheromone of the southern pine beetle <u>Dendroctonus frontalis</u>. Although the synthesis of this compound in optically active form has been accomplished by the derivation of either optically active natural products, 2,3 or optically resolved compounds, the reports on its asymmetric synthesis are rare. One is the work of Mukaiyama in which the diastereoselective alkylation of keto aminal is utilized as the key step. The other is the procedures of Lee and Scharf where the chiral center is constructed by the enantioselective epoxidation of allylic alcohols. Described herein is the shortest route to (+)- and (-)- frontalin with readily available reagents.

The straightforward, three-step synthesis of (S)-(-)-frontalin ($\frac{1}{2}$) is illustrated in Scheme 1. β -Methallyl alcohol ($\frac{2}{2}$) was enantioselectively epoxidized by the procedure of Katsuki-Sharpless (2 equiv. of t-BuOOH, CH₂Cl₂,

Scheme 1.

4~

1530 Chemistry Letters, 1985

-20 °C) using (D)-(-)-diethyl tartrate {Aldrich, $[\alpha]_D^{25}$ -8.5 (neat)} in 44% yield. The epoxy alcohol $\frac{3}{3}$ { $[\alpha]_D^{25}$ +10.7 (CHCl $_3$)} thus obtained was converted into olefinic diol $\frac{4}{2}$ in 85% yield upon treatment with homoallylmagnesium bromide in the presence of 0.1 equiv. of Li_2CuCl_4 (THF, -78 °C). The intramolecular acetalization of $\frac{1}{2}$ of $\frac{4}{4}$ catalyzed by PdCl $_2$ (0.1 equiv.) in the presence of CuCl (1 equiv.) and $\frac{1}{2}$ and $\frac{1}{2}$ in triglyme at 50 °C gave a 92% ee of (-)-frontalin { $[\alpha]_D^{23}$ -50.9 (ether)} in 76% isolated yield. In this acetalization, the use of triglyme as the solvent allowed us to make the isolation of volatile frontalin easier. $\frac{12}{2}$

According to the same sequence described in Scheme 1, unnatural (R)-(+)-frontalin was obtained in 83% ee from (-)-epoxy alcohol prepared by the asymmetric epoxidation of 2 with (L)-(+)-diethyl tartrate {Tokyo Kasei, $[\alpha]_D^{20}$ +7.6 (neat)}. Optically active pheromones such as brevicomin and multistriatin are constituted of the same structural unit of bicyclic acetal as that of frontalin. Our approach outlined here provides a high potentiality for the asymmetric synthesis of these insect pheromones.

References

- 1) G. W. Kinzer, A. F. Fentiman, Jr., T. F. Page, Jr., R. L. Foltz, J. P. Vite, and G. B. Pitman, Nature, 221, 477 (1969).
- 2) R. Barner and J. Hübscher, Helv. Chim. Acta, <u>66</u>, 880 (1983).
- 3) R. Naef and D. Seebach, Liebigs Ann. Chem., 1983, 1930 and references cited therein.
- 4) K. Mori, Tetrahedron, 31, 1381 (1975).
- 5) Y. Sakito and T. Mukaiyama, Chem. Lett., 1979, 1027.
- 6) A. W. M. Lee, J. Chem. Soc., Chem. Commun., 1984, 578.
- 7) C. Meister and H.-D. Scharf, Liebigs Ann. Chem., 1983, 913.
- 8) The synthesis of (-)-frontalin utilizing the baker's yeast-mediated transformation of α -methylcinnamaldehyde has been also reported; see, C. Fuganti, P. Grasselli, and S. Servi, J. Chem. Soc., Perkin Trans. 1, <u>1983</u>, 241.
- 9) T. Katsuki and K. B. Sharpless, J. Am. Chem. Soc., 102, 5976 (1980).
- 10) M. H. Brookes, B. T. Golding, D. A. Howse, and A. T. Hudson, J. Chem. Soc., Chem. Commun., 1983, 1051.
- 11) T. Hosokawa, T. Ohta, and S.-I. Murahashi, J. Chem. Soc., Chem. Commun., 1983, 848; N. T. Byrom, R. Grigg, B. Kongkathip, G. Reimer, and A. R. Wada, J. Chem. Soc., Perkin Trans. 1, 1984, 1643.
- 12) The use of other solvents such as DME, acetone, and diglyme also afforded frontalin quantitatively. However, because of lower boiling points of these solvents, it was isolated only in quite low yield (≈20%).
- 13) R. M. Silverstein, R. G. Brownlee, T. E. Bellas, D. L. Wood, and L. E. Browne, Science, 159, 889 (1968).
- 14) G. T. Pearce, W. E. Gore, R. M. Silverstein, J. W. Peacock, R. A. Cuthbert, G. N. Lanier, and J. B. Simeone, J. Chem. Ecol., 1, 115 (1975).

(Received July 26, 1985)